

# Interview Questions

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## 40 TOP Stoichiometry Online Test – Multiple Choice Questions and Answers

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Congratulations – you have completed **Stoichiometry Online Test Set 1**.

You scored **6** points out of **10** points total.

Your answers are shown below:

1. Sometimes, in chemical processes, a part of the outlet stream is rejected as waste in order to keep the impurity level in the system within limits. This phenomenon is termed as the

- **recycling**

NFW

- purging
- bypassing
- recirculation

2. 2 litres of nitrogen at N.T.P. weighs \_\_\_\_\_ gms

- 1.25
- **2.5**
- 14
- 28

3. The equilibria relations in a multicomponent and multiphase system can not be calculated with the help of the

- experimental data
- empirical equations
- **phase rule**
- theoretical equations

4. What is the total pressure exerted by a mixture of 0.45 kg mole of benzene, 0.44 kg mole of toluene and 0.23 kg mole of o-xylene at 100°C, if their vapor pressures at 100°C are 1340, 560 and 210 mmHg respectively ?

- 756.2
- **780.5**
- 801.5
- 880.5

5. In case of vapor-liquid equilibria, which of the following does not account for gas phase deviation from ideality?

- Activity co-efficient
- Fugacity co-efficient
- **Use of equation of state**
- None of these

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6. A vapor whose partial pressure is less than its equilibrium vapor pressure is called a \_\_\_\_\_ vapor

- supersaturated
- saturated
- **superheated**
- none of these

7. In case of a solution (not of a solid in a liquid), whose total volume is more than the sum of volumes of its components in their pure states, solubility is

- unchanged by the pressure change
- **increased with the increase in pressure**
- decreased with the increase in pressure
- independent of the temperature

8. An azeotropic solution of two liquids has boiling point lower than either of them, when it

- is saturated
- is unsaturated
- **shows positive deviation from Raoult's law**
- shows negative deviation from Raoult's law

9. If the pressure of a gas is reduced to half & its absolute temperature is doubled, then the volume of the gas will

- increase four times
- be reduced to 1/4th
- **increase two times**
- none of these

10. Pick out the wrong statement:

- A vapor whose partial pressure is less than its equilibrium vapor pressure, is termed as a 'superheated vapor'

- **The temperature at which a vapor is saturated is termed as the boiling point**
- A vapor is termed as a saturated vapor, if its partial pressure equals its equilibrium vapor pressure
- The difference between the existing temperature of a vapor and its saturation temperature (i.e. dew point) is called its 'degree of superheat'

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